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Reading: | • 74 L On15L 0-59L Ending
Buret Reading: 11.83 L10-10L 10.591 mL of
Titrant Used: 10.09L/9.95L10.00L mmole of
NaOH: 1.009 moll 9.95 no 1.0 ml mmole of
H₂SO₄ 0.5 mal 5.0 0.5 5.0nd Molarity of
H₂SO₄ Solution 0.05 M 0.5 M 0.05 M

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0.05M Average Molarity 0.6M.

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Reaction 1: $1 \text{ NaHCO}_3 (\text{s}) + 1 \text{ NaOH} (\text{s}) + 1 \text{ CO}_2 (\text{g})$
Reaction 2: $2 \text{ NaHCO}_3 (\text{s}) + 1 \text{ Na}_2\text{CO}_3 (\text{s}) + 1 \text{ CO}_2 (\text{g}) + 1 \text{ H}_2\text{O} (\text{g})$

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Reaction 3: $2 \text{NaHCO}_3 (\text{s}) + 1 \text{Na}_2\text{O} (\text{s}) +$
2. Propose a voltaic cell that would coat a piece of copper with platinum. Include the half reactions for your cell and the voltage potential based on the.

CHEM 103 - UIUC

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1. Identify the group for each unknown.
Your unknowns will be from three different
groups. (Example: A K^+ salt, a halogen, and
an acid.) 2. Identify the specific compound
from each group. (Example: KCl , Br_2 and
 HCl) 3. Explain the reasoning that enabled
you to come to these conclusions.

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Lab Book: Chemistry 104 Laboratory Manual, Fall 2014 Chemistry Department, University of Wisconsin-Madison; available in the chemistry building lobby from Alpha Chi Sigma, cash only. Lab Notebook: Carbonless laboratory notebook with duplicate pages available from Alpha Chi Sigma or local bookstores.

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