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~~ANSWERS,,~~

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~~QUESTIONS~~ *Metals*

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with Formulas
and Equations:~~

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Review. a system of physical units (SI Units) based on the meter, kilogram, second, ampere, kelvin, candela, and mole, together with a set of prefixes to indicate multiplication

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or division by a
power of ten.

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all the
questions in

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complete sentences (just restate the original question). Show the math work for # 27-30. 1. The law of conservation of mass follows from the concept that _____. 2. If 4 g of element A

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combine with 10
g of element B,
then 12 g of
element A
combine with
_____ g of
element B. 3.

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five principles
of the atomic...

law of

conservation of
mass. law of

multiple

proportions. law
of definite

proportions. 1.

all matter is

composed of

extremely small

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particles called... mass is
neither created
nor destroyed in
a chemical
reaction b....

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that has mass
and takes up
space. matter
composed of all
one type of atom
(can be a
single element,
like He or a
diatomic
element, like O₂
or H₂). Elements
cannot be broken
down into
simpler

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chemical
combinations of
two or more
elements.

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The presentation of information in this chapter assumes that you can already perform the tasks listed below. You can test your readiness to proceed by answering the Review .

Questions at the

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end of the chapter. This might also be a good time to read the Chapter . Objectives, which precede the Review Questions.

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particles.

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Matter is made
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approaching the
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law of conservation of mass. mass is neither created nor destroyed during ordinary chemical reactions or physical changes
 $H_2 + O = H_2 O$. law of definite proportions.

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of definite

proportions.

compounds are

made of the same

elements in the

same

proportions. law

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of conservation
of mass. mass
cannot be

created or

destroyed. law

of multiple

proportions.

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Answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____.

17. The energy required to remove one

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electron from an
atom is called
its _____. 18.

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