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~~Part 1 (4th of 19 Chapters) Elements~~  
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Kinetic Theory Of Matter \u0026amp;

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Mixtures? | Chemistry Matters Chemistry  
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Speed of Reaction

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Atomic Structure Part 1

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17.3 17.4 Activation Energy, Energy

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(03A Movement of Substances: Part1)

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Private Limited 16.3 (c) The nuclear disaster has caused radioactive substances to spread into the atmosphere and pollute the air.

Chapter 16 The Periodic Table - O Level

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The error in print is on Page 271 under the Chapter of “ Electrolysis ” in Chemistry Matters for Pure Chemistry syllabus. In the Key Ideas which acts like sub-section summary, it states that “ OH- ions will be discharged as oxygen gas. However, if a solution is concentrated in halide ions, the halogen gas will be

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9.2 (c) Number of moles of  $\text{C l}_2$  gas = mass  
of  $\text{C l}_2$  / molar mass of  $\text{C l}_2$  =  $1.4271 / 71 =$   
 $0.02$  mol (d) Number of moles of sodium  
chloride = number of ions of sodium  
chloride  $6 \times 10^{23} = 6 \times 10^{22}$   $6 \times 10^{23} = 0.1$   
mol  $2$ .

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Test Yourself 4.1 (page 74) 1. (Any four differences) Mass Weight • An amount of matter • A gravitational force • A scalar quantity (i.e. has only magnitude) • A vector quantity (i.e. has both magnitude and direction) • SI unit: kilogram (kg) • SI

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10.2 Test Yourself 10.2 (page 167) 1.

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$2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$   
Volume of carbon dioxide formed = 2  
x

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